Number	AL	
1 various	COS#	
1.	#1.0	Classify sodium chloride as an element, mixture, compound, or colloid.
2.	#1.0	Classify air as an element, mixture, compound, or colloid.
3.	#1.0	Classify a blueberry muffin as an element, mixture, compound, or colloid.
4.	#1.0	Classify calcium carbonate as an element, mixture, compound, or colloid.
5.	#1.0	Classify aluminum as an element, mixture, compound, or colloid.
6.	#1.0	Classify carbon as an element, mixture, compound, or colloid.
7.	#1.0	Classify salt water as an element, mixture, compound, or colloid.
8.	#1.0	Classify copper as an element, mixture, compound, or colloid.
9.	#1.0	Classify carbon dioxide as an element, mixture, compound, or colloid.
10.	#1.0	Classify muddy water as an element, mixture, compound, or colloid.
11.	#1.1	Is mass an example of an intensive property or an extensive property?
12.	#1.1	Is boiling point an example of an intensive property or an extensive property?
13.	#1.1	Is density an example of an intensive property or an extensive property?
14.	#1.1	Is weight an example of an intensive property or an extensive property?
15.	#1.1	Is height an example of an intensive property or an extensive property?
16.	#1.1	Is melting point an example of an intensive property or an extensive property?
17.	#1.1	Is freezing point an example of an intensive property or an extensive property?
18.	#1.1	Is width an example of an intensive property or an extensive property?
19.	#1.1	Is volume an example of an intensive property or an extensive property?
20. 21.	#1.1	Is malleability an example of an intensive property or an extensive property?
22.	#1.2	To what category of elements does an element belong if it is a poor conductor of electricity?  To what category of elements does an element belong if it is a good conductor of electricity?
23.	#1.2	To what category of elements does an element belong if it is a good conductor of electricity?  To what category of elements does an element belong if it is malleable?
24.	#1.2	To what category of elements does an element belong if it is brittle?
25.	#1.2	To what category of elements does an element belong if it is office?  To what category of elements does an element belong if it tends to gain electrons when forming an ion?
26.	#1.2	To what category of elements does an element belong if it tends to lose electrons when forming an ion?
27.	#1.2	To what category of elements does an element belong if it is ductile?
28.	#1.2	To what category of elements does an element belong if it is duetile?  To what category of elements does an element belong if it exhibits luster?
29.	#1.2	To what category of elements does an element belong if it is dull?
30.	#1.2	To what category of elements does an element belong if it is a good conductor of heat?
31.	#1.3	What type of mixture is sand and water?
32.	#1.3	What type of mixture is black coffee?
33.	#1.3	What type of mixture is oil and water?
34.	#1.3	What type of mixture is blood?
35.	#1.3	What type of mixture is a chocolate chip cookie?
36.	#1.3	What type of mixture is a blueberry muffin?
37.	#1.3	What type of mixture is chicken noodle soup?
38.	#1.3	What type of mixture is air?
39.	#1.3	What type of mixture is dirt?
40.	#1.3	What type of mixture is concrete?
41.	#3.0	Given: Li, Rb, K, or Na, which has the largest atomic radius?
42.	#3.0	Given: Li, Rb, K, or Na, which has the smallest atomic radius?
43.	#3.0	Given: K, Cs, As, Br, which has the largest atomic radius?
44.	#3.0	Given: K, Cs, As, Br, which has the smallest atomic radius?
45.	#3.0	Given: K, Cs, As, Br, which has the largest electronegativity?
46.	#3.0	Given: K, Cs, As, Br, which has the smallest electronegativity?
47.	#3.0	Given: Li, Rb, K, Na, which has the smallest electronegativity?
48.	#3.0	Given: Li, Rb, K, Na, which has the largest electronegativity?
49.	#3.0	How many valence electrons does phosphorus have?

50. #3.0 (Given: iodine, helium, sulfur, or hydrogen, which is a halogen?  51. #3.0 (How many valence electrons does Si have?)  52. #3.0 (How many valence electrons does Rn have?)  53. #3.0 (What factor determines the reactivity of an element?)  54. #3.1 (Given an orbital diagram be able to determine the element.)  55. #3.1 (Given an orbital diagram be able to determine the element.)  56. #3.1 (Given an orbital diagram be able to determine the element.)  57. #3.1 (What is the correct electron dot structure for sodium?)  58. #3.1 (How many electrons can a single orbital hold?)  59. #3.1 (How many electrons can a single orbital hold?)  59. #3.1 (Which element 's electron configuration ends in 64°?)  60. #3.1 (Which element selectron configuration ends in 64°?)  61. #3.1 (Which element selectron configuration ends in 64°?)  62. #3.1 (Which holbe gas has ONLY two dots in its Lewis dot structure?)  63. #3.1 (What is the correct electron dot structure for oxygen?)  64. #3.1 (What is the correct electron dot structure for oxygen?)  65. #3.2 (Calculate the number of protons in borno.)  66. #3.2 (Calculate the number of protons in borno.)  66. #3.2 (Calculate the number of protons in borno.)  67. #3.2 (Calculate the number of protons in potassium-40.  68. #3.2 (Calculate the number of protons in potassium-40.  69. #3.2 (Calculate the number of protons in potassium-40.  70. #3.2 (Calculate the number of protons in oxygen-17.  71. #3.2 (Calculate the number of protons in oxygen-17.  72. #3.2 (Calculate the number of protons in oxygen-17.  73. #3.2 (Calculate the number of protons in 3°P.  74. #3.2 (Calculate the number of protons in 3°P.  75. #3.2 (Calculate the number of protons in 3°P.  76. #3.2 (Calculate the number of protons in 3°P.  77. #3.2 (Calculate the number of protons in 3°P.  78. #3.2 (Calculate the number of protons in 3°P.  79. #3.2 (Calculate the number of protons in 3°P.  79. #3.2 (Calculate the number of protons in 3°P.  79. #3.3 (Which scientist came up with the atomic theory of matter?)  80. #3.3 (Whic	50	112.0	
53. #3.0   How many valence electrons does Rn have?			
53. #3.0 What factor determines the reactivity of an element?  54. #3.1 Given: vanandium, bromine, strontium, or neodymium, in which would you find f orbitals being filled?  55. #3.1 Given an orbital diagram be able to determine the element.  56. #3.1 What is the correct electron dot structure for sodium?  57. #3.1 What is the correct electron dot structure for sodium?  58. #3.1 How many electron can a single orbital hold?  59. #3.1 Which element's electron configuration ends with 49.?  60. #3.1 Which element's electron configuration ends with 49.?  61. #3.1 Which element's electron configuration ends with 49.?  62. #3.1 Which oble gas has ONA! Yew dots in its Lewis dot structure?  63. #3.1 What is the correct electron dot structure for oxygen?  64. #3.1 What is the correct electron dot structure for oxygen?  65. #3.2 Calculate the number of protons in boron.  66. #3.2 Calculate the number of protons in magnesium.  67. #3.2 Calculate the number of protons in potassium-40.  68. #3.3 Calculate the number of neutrons in patassium-40.  69. #3.2 Calculate the number of protons in potassium-40.  70. #3.2 Calculate the number of protons in potassium-40.  71. #3.2 Calculate the number of protons in oxygen-17.  72. #3.3 Calculate the number of protons in oxygen-17.  73. #3.3 Calculate the number of protons in oxygen-17.  74. #3.3 Calculate the number of protons in 325 U.  75. #3.2 Calculate the number of protons in 325 U.  76. #3.2 Calculate the number of neutrons in 326 U.  77. #3.3 Calculate the number of protons in 327 U.  78. #3.3 Calculate the number of protons in 327 U.  79. #3.3 Calculate the number of protons in 327 U.  71. #3.3 Calculate the number of protons in 327 U.  72. #3.3 Calculate the number of protons in 328 U.  73. #3.3 Calculate the number of protons in 328 U.  74. #3.3 Calculate the number of protons in 328 U.  75. #3.3 Calculate the number of protons in 328 U.  76. #3.3 Calculate the number of protons in 328 U.  77. #3.3 Calculate the number of protons in 328 U.  78. #3.3 Which isotope of cono cont			,
34.   #3.1   Given vanadium, bromine, strontium, or neodymium, in which would you find f orbitals being filled?   55.   #3.1   Given an orbital diagram be able to determine the element.   56.   #3.1   Given an orbital diagram be able to determine the element.   57.   #3.1   What is the correct electron dot structure for sodium?   58.   #3.1   How many electrons can a single orbital hold?   59.   #3.1   Which element's electron configuration ends with 4p²?     60.   #3.1   Which element's electron configuration ends with 4p²?     61.   #3.1   Which element shape as configuration of   Ne[3s³g¹?			·
55. #3.1 Given an orbital diagram be able to determine the element.			, and the second
56, #3.1   Given an electron configuration determine what element is represented.			
St. #3.1   What is the correct electron dot structure for sodium?			
Section			
59, #3.1   Which element's electron configuration ends with \$\frac{4}{9}\$?   60. #3.1   Which element's electron configuration of \$\text{Index} \frac{7}{2}\$?   61. #3.1   Which element's electron configuration of \$\text{Index} \frac{7}{2}\$?   62. #3.1   Which noble gas has \$\text{ONLY}\$ two dots in its Lewis dot structure?   63. #3.1   What is the correct electron dot structure for structure?   64. #3.1   What is the correct electron dot structure for strontium?   65. #3.2   Calculate the number of protons in boron.   66. #3.2   Calculate the number of electrons in magnesium.   67. #3.2   Calculate the number of neutrons in carbon = 14.   68. #3.2   Calculate the number of protons in potassium=40.   69. #3.2   Calculate the number of protons in potassium=40.   70. #3.2   Calculate the number of protons in carbon = 17.   71. #3.2   Calculate the number of protons in carbon = 17.   72. #3.2   Calculate the number of protons in carbon = 17.   73. #3.2   Calculate the number of protons in oxygen = 17.   74. #3.2   Calculate the number of neutrons in \$\frac{19}{2}\$.   75. #3.2   Calculate the number of neutrons in \$\frac{19}{2}\$.   75. #3.2   Calculate the number of protons in \$\frac{135}{2}\$.   76. #3.2   Calculate the number of protons in \$\frac{135}{2}\$.   77. #3.2   Calculate the number of protons in \$\frac{135}{2}\$.   78. #3.2   Calculate the number of protons in \$\frac{135}{2}\$.   79. #3.2   Calculate the number of protons in \$\frac{135}{2}\$.   79. #3.2   Calculate the number of protons in \$\frac{135}{2}\$.   79. #3.2   Calculate the number of protons in \$\frac{135}{2}\$.   79. #3.2   Calculate the number of protons in \$\frac{135}{2}\$.   79. #3.2   Calculate the number of protons in \$\frac{135}{2}\$.   79. #3.3   Which isotope of oxygen contains the same number of neutrons and protons?   80. #3.3   Which isotope of oxygen contains the same number of neutrons and protons?   81. #3.3   Which isotope of oxygen contains the same number of neutrons and protons?   82. #3.3   Which isotope of oxygen contains the same			
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101.	#5.0	What happens to the distance between molecules of a substance when it vaporizes?
102.	#5.0	How does the kinetic energy of a hydrogen molecule at 25°C compare to the kinetic energy of an
		oxygen molecule at 25°C?
103.	#6.1	Predict the type of bond that is formed in the diagram above.
		$Na^+ + Cl^- \rightarrow NaCl$
104.	#6.1	Predict the type of bond that is formed in the diagram above.
		$Cl + Cl \rightarrow Cl_2$
105.	#6.1	Predict the type of bond that would form between lithium and fluorine.
106.	#6.1	24
100.	#6.1	Predict the type of bond that would form between sodium and fluorine.  Predict the type of bond that would form between carbon and hydrogen.
107.	#6.1	Predict the type of bond that would form between capper and chlorine.
108.	#6.1	Predict the type of bond formed by the sharing of four electrons.
110.	#6.1	Which diatomic molecule forms a triple covalent bond?
110.	$\pi 0.1$	which diatomic molecule forms a triple covalent bond:
111.	#6.1	Predict the type of bond that would form between sulfur and oxygen.
112.	#6.2	Group 2A elements, alkaline earth metals, tend to form what charge?
113.	#6.2	What would the oxidation number be for the element with the following electron configuration:
		$1s^22s^22p^63s^23p^5$ ?
114.	#6.2	What would be the oxidation number for the following element: $1s^2 2s^2 2p^5$ ?
115.	#6.2	What would be the oxidation number for the following element: $[Ar] 4s^2$ ?
116.	#6.2	Given: nitrogen, sulfur, boron, or chlorine, which would tend to form an ion with a –3 charge?
117.	#6.2	Given: carbon, oxygen, barium, or potassium, which would tend to form an ion with a +2 charge?
118.	#6.2	What is the sum of the oxidation numbers in a neutral compound?
119.	#6.2	What is the charge of the phosphate ion in $K_3PO_4$ ?
120.	#6.2	What is the oxidation number of hydrogen in H <sub>2</sub> O?
121.	#6.2	What is the oxidation number of the chromate ion in potassium chromate?
122.	#6.2	What is the oxidation number of phosphorus in sodium phosphate?
123.	#6.3	What is the name of $H_2SO_4$ (aq)?
124.	#6.3	What is the formula for aluminum sulfate?
125.	#6.3	What is the formula for copper (II) chloride?
126.	#6.3	What is the correct formula for carbon disulfide?
127.	#6.3	What is the correct name for CCl <sub>4</sub> ?
128.	#6.3	What is the correct name for MgSO <sub>3</sub> ?
129.	#6.3	What is the formula of calcium phosphate, which is made up of the ions Ca <sup>2+</sup> and PO <sub>4</sub> <sup>3-</sup> ?
130.	#6.3	What is the correct name for $(NH_4)_3PO_4$ ?
131.	#6.3	What is the correct name for Li <sub>2</sub> CO <sub>3</sub> ?
132.	#6.3	What is the correct name for NaOH?
133.	#6.3	What is the correct formula for potassium permanganate?
134.	#6.3	What is the correct formula for hydrochloric acid?
135.	#6.3	What is the correct formula for phosphoric acid.?
136.	#8.0	Is burning wood an example of a chemical change or a physical change?
137.	#8.0	Is freezing water an example of a chemical change or a physical change?
138.	#8.0	Is malleability an example of a chemical property or a physical property?
139.	#8.0	Is melting point an example of a chemical property or a physical property?
140.		Is "ignites when dropped into water" an example of a chemical property or a physical property?
141. 142.	#8.0	Is conducting electricity an example of a chemical property or a physical property?  Is an object's tendency to rust an example of a chemical property or a physical property?
142.	#8.0	Is an object's tendency to fust an example of a chemical property or a physical property?  Is an object's tendency to tarnish an example of a chemical property or a physical property?
143.	#8.0	Is an object's ductility an example of a chemical property or a physical property?
144.	#8.0	Is an object's luster an example of a chemical property or a physical property?
173.	11 U.U	15 an object 5 laster an example of a chemical property of a physical property:

146.	#9.3	Identify the three main types of radioactivity from the list below.
147.	#9.3	Determine what type of radiation is being given off from a nuclear transition equation.
148.	#9.3	What particle makes up beta radiation?
149.	#9.3	In the equation $^{238}U \rightarrow ^{234}$ Th, what type of radiation is given off?
150.	#9.3	In the equation $^{14}C \rightarrow ^{14}N$ , what type of radiation is given off?
151.	#9.3	What type of radiation is not deflected by electric or magnetic fields?
152.	#9.3	What is the product of all nuclear fission reactions?
153.	#9.3	What condition is required for fusion reactions to occur?
154.	#9.3	What particle does Argon – 39 lose when it decays to potassium – 39?
155.	#9.3	What is a beta particle?
156.	#10.0	What is a control?
157.	#10.0	If you know an object's density, what else do you need to know in order to calculate its mass?
158.	#10.0	What is the density of an object with a mass of 7.5 grams and a volume of 5.0 mL?
159.	#10.0	Which prefix is equivalent to 1000 or 10 <sup>3</sup> ?
160.	#10.0	Which prefix is equivalent to 0.001 or 10 <sup>-3</sup> ?
161.	#10.0	What is the SI base unit used to measure the amount of a substance?
162.	#10.0	What is the SI base unit for temperature?
163.	#10.0	According to the rules for significant digits, how many significant digits will be in the answer to the
		problem $23.32 \times 6.59$ ?
164.	#10.0	Solve the problem 36.28 + 45.7 using the correct number of significant digits.
165.	#10.0	What is the number <b>1,592,000,000</b> when written in proper scientific notation?
166.	#10.0	What is the number 3.55 x 10 <sup>-3</sup> when written in ordinary notation?