

Name: _____

Write the molecular, complete ionic, & net ionic equations for all of the following reactions. Circle or highlight the net ionic equation as your answer. This will need to be done on a separate sheet of paper. Attach all of your work to this sheet.

1. Solutions of sodium hydroxide and acetic acid are mixed.
2. Hydrogen sulfide gas is bubbled through excess potassium hydroxide solution.
3. Solid barium oxide is added to distilled water.
4. Solid calcium oxide is exposed to a stream of carbon dioxide gas.
5. Solid dinitrogen pentoxide is added to water.
6. Lithium metal is burned in air.
7. Solid ammonium carbonate is heated.
8. A solution of hydrogen peroxide is catalytically decomposed.
9. A solution of ammonium sulfate is added to a saturated solution of barium hydroxide.
10. A solution of copper(II) chloride is added to a solution of sodium sulfide.
11. Solutions of manganese(II) sulfate and ammonium sulfide are mixed.
12. Aluminum metal is added to a solution of copper(II) chloride.
13. Magnesium metal is added to nitrogen gas.
14. Solid lithium hydride is added to distilled water.
15. Solid lithium oxide is added to excess water.
16. Solid potassium chlorate is heated in the presence of a catalyst.
17. Dilute hydrochloric acid is added to a solution of potassium sulfite.
18. A mixture of solid calcium oxide and solid tetraphosphorus decaoxide is heated.
19. Sulfur dioxide gas is passed over solid calcium oxide.
20. Hexane is combusted
21. Sulfuric acid reacts with sodium carbonate
22. Nonyne is combusted
23. Phosphoric acid decomposes
24. Solid magnesium chloride reacts with oxygen
25. Barium hydroxide decomposes